

Hill-Petrucci  
Homework  
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23. Noble Gas electron configurations: a,b,c,e

$$\text{Not: } \text{Zn}^{2+} = 1s^2 2s^2 2p^6 3s^2 3p^6 3d^{10}$$

$$\text{Hg}^{2+} = [\text{Xe}] \quad 4f^{14} 5d^{10}$$

24. a.  $[\text{Ar}]4s^2 3d^{10} 4p^6$

b.  $[\text{Ne}]3s^2 3p^6$

c.  $[\text{Ar}]4s^2 3d^{10} 4p^6$

d.  $[\text{Kr}]4d^{10}$

e.  $[\text{Ar}]4s^2 3d^{10} 4p^6$

f.  $[\text{He}]2s^2 2p^6$

